

## 1) Ionic Compounds – Formula

Name	Cation	Anion	Formula
Sodium chloride			
Magnesium chloride			
Calcium oxide			
Lithium phosphide			
Aluminum sulfide			
Calcium nitride			

## 2) Ionic Compounds – Names

Formula	Cation	Anion	Name
K <sub>2</sub> S			
BaF <sub>2</sub>			
MgO			
Na <sub>3</sub> N			
AlCl <sub>3</sub>			
Mg <sub>3</sub> P <sub>2</sub>			

### 3) Ionic Compounds (Transition Metals) – Formula

Name	Cation	Anion	Formula
Iron (III) chloride			
Iron (II) oxide			
Copper (I) sulfide			
Copper (II) nitride			
Zinc oxide			
Silver sulfide			

### 4) Ionic Compounds (Transition Metals) – Name

Formula	Cation	Anion	Name
$\text{Cu}_2\text{S}$			
$\text{Fe}_2\text{O}_3$			
$\text{CuCl}_2$			
$\text{FeS}$			
$\text{Ag}_2\text{O}$			
$\text{FeBr}_2$			

5) Ionic Compounds (Polyatomic Ions) – Formula

Name	Cation	Anion	Formula
Potassium carbonate			
Sodium nitrate			
Calcium hydrogen carbonate			
Manganese (II) sulfate			
Chromium (III) nitrite			
Lithium phosphate			

6) Ionic Compounds (Polyatomic Ions) – Name

Formula	Cation	Anion	Name
CaSO <sub>4</sub>			
Al(NO <sub>3</sub> ) <sub>3</sub>			
Na <sub>2</sub> CO <sub>3</sub>			
MgSO <sub>3</sub>			
Cu(OH) <sub>2</sub>			
Mg <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>			

### 7) Covalent Compounds – Formula

Name	Formula
Sulfur trioxide	
Dinitrogen monoxide	
Iodine pentachloride	
Diphosphorous pentaoxide	
Dihydrogen sulfide	
Sulfur difluoride	

### 8) Covalent Compounds – Names

Formula	Name
$N_2O_5$	
$PCl_3$	
$BrCl_3$	
$CS_2$	
$BF_3$	
$XeI_2$	