Ch2/Worksheet/Nomenclature **Name:**

1) Write the ions that can form from the following elements and the names. An example is given in the first row.

Na^+	Sodium ion	iodide	I^-
Fe^{2+}			P^{3-}
	Calcium ion	Oxide	
Cr^{6+}		Hydride	

2) Write the names or give the formulas of the following polyatomic ions.

OH^-			Carbonate
	Phosphate		
NO_3^-			Hydroxide
	Ammonium ion	SO_4^{2-}	

3) Classify the following as covalent or ionic compounds

P_2O_5	Iron (II) Oxide	Sodium oxide
Calcium chloride	NO_3	CO_2
BaO	$AlCl_3$	MnS_2

4) Give four diatomic compounds with the same name as their element name.

5) Identify and name the polyatomic ions in the following compounds.

Na_2SO_4	KNO_3	NH_4Cl
$Ca_3(PO_4)_2$	NH_4NO_3	$Sr(OH)_2$

6) Give the names or formulas of the following compounds.

K_3N		silver bromide	
SO_2		silicon dioxide	
$Pb(SO_4)_2$		carbon tetrachloride	
$Fe(NO_3)_3$		lead (II) nitride	
$Al(CN)_3$		tin (II) nitrite	
$Mn_2(SO_3)_3$		cobalt (III) oxide	
$SnSe_2$		chromium (III) hydroxide	
$Be(HCO_3)_2$		titanium (II) acetate	
CuOH		magnesium sulfate heptahydrate	
NH_4Cl		potassium carbonate	
Cu_3P		diboron tetrabromide	
$Ca(C_2H_3O_2)_2$		lithium iodide	
$FePO_4$		silver acetate	
NaBr		manganese (II) phosphate	
P_2O_5		chromium (VI) phosphate	
$Zn(NO_2)_2$		vanadium (V) sulfide	
		nickel (III) sulfide	