

Ch4/Worksheet/Solution Electrolytes-Ppt Equations **Name:** _____

- 1) Identify the following substances as strong, weak or non electrolytes.
- a) HCl b) NH₃ c) C₆H₁₂O₆ (glucose) d) N₂
- e) KCl f) CaF₂ g) HNO₃ h) NaOH
- 2) Circle the salts that will be soluble in water.
- a) AgBr b) AgNO₃ c) K₂SO₄
- d) CuS e) Mg(OH)₂ f) PbSO₄
- 3) Indicate whether a precipitate will form when the following solutions are mixed. Write the molecular equation.
- a) iron(III) nitrate and potassium hydroxide
- b) sodium sulfide and nickel(II) sulfate

- 4) Write the products and all the equations in the table below:

Molecular Equation	$\text{AgNO}_3(aq) + \text{K}_2\text{Cr}_2\text{O}_7(aq) \longrightarrow$
Total Ionic Equation	
Net Ionic Equation	

5) Write the complete reaction between the following compounds. Write the molecular, ionic and net ionic equation for all the reactions.

a) Lead (II) nitrate and sodium chloride

b) ammonium chloride and lithium carbonate