

1. Classify each of the following substances as: 1) acid, base, or neutral, and 2) strong or weak.
a) HNO_3 b) HClO c) NH_3 d) NaNO_3 e) Ba(OH)_2
2. Complete the following neutralization equations by writing the products and all the equations.

Molecular Equation	$\text{Mg(OH)}_2(\text{s}) + \text{HCl}(\text{aq}) \longrightarrow$
Total Ionic Equation	
Net Ionic Equation	

Molecular Equation	$\text{NaOH}(\text{aq}) + \text{H}_2\text{SO}_4(\text{aq}) \longrightarrow$
Total Ionic Equation	
Net Ionic Equation	