- 1. Write the oxidation numbers of the underlined element in the following compounds.
 - a) <u>Cr</u>Cl₃
- b) <u>Mn</u>O₄
- c) K₂<u>Cr</u>O₄
- d) <u>C</u>H₄
- 2. Write the oxidation numbers of the elements underlined in the following equations. Indicate which one is getting oxidized and which one is getting reduced. Also write which is the oxidizing agent in the reactions.
 - a) HCl + <u>O</u>2
- \longrightarrow Cl₂ + H₂O

b) \underline{Ag} + \underline{H}^+ + $\underline{NO_3}^ \longrightarrow$ \underline{Ag}^+ + $\underline{H_2O}$ + \underline{NO}

c) $\underline{Se}O_3^{2-} + \underline{I}^- + \underline{H}^+ \longrightarrow \underline{Se} + \underline{I}_2 + \underline{H}_2O$