

1) Complete the following table.

Molecule	(1) $\underline{\text{SO}}_2$	(2) $\underline{\text{HBF}}_2$	(3) $\underline{\text{XeF}}_4$	(4) $\underline{\text{CH}}_2\text{Cl}_2$	(5) $\underline{\text{NF}}_3$
Lewis structure					
Electron groups on central atom					
AXE formula					
Name of electron geometry on the central atom					
Hybridization					
Bond angle(s)					
Shape of molecule					

2) Predict a) the bond angle, (b) the hybridization around the indicated atoms (the atoms to which the arrows are drawn in the structures below). Write your answers near the corresponding labels (1 to 5) in the drawings. (Note: the lone pairs on the F atoms are omitted.)

