

1) Complete the following table.

Molecule	(1) $\underline{\text{S}}\text{O}_2$	(2) $\text{H}\underline{\text{B}}\text{F}_2$	(3) $\underline{\text{X}}\text{eF}_4$	(4) $\underline{\text{C}}\text{H}_2\text{Cl}_2$	(5) $\underline{\text{N}}\text{F}_3$
a) Lewis structure					
b) Electron groups on central atom.					
c) AXE formula					
d) Name of geometry on the central atom					
e) Bond angle associated with the geometry					
f) Name of shape of the molecule					
g) Hybridization					

2) Predict a) the bond angle, (b) the hybridization around the indicated atoms (the atoms to which the arrows are drawn in the structures below). Write your answers near the corresponding labels (1 to 5) in the drawings. (Note: the lone pairs on the F atoms are omitted.)

