

Electrolytes - 1

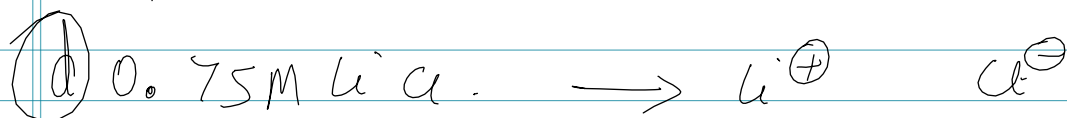
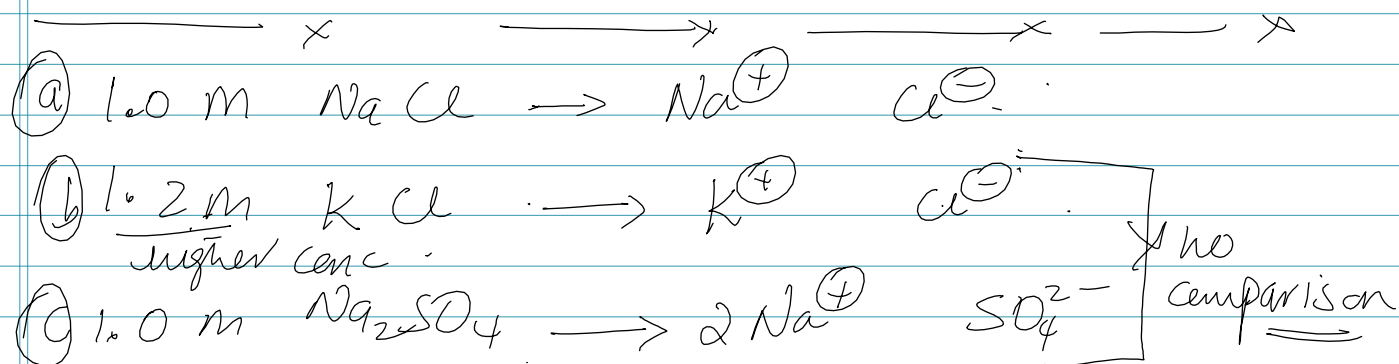
Identify the following as strong, weak or nonelectrolytes.

(a) $\text{NaBr} \rightarrow$ ionic - dissociated in H_2O
Strong electrolyte.

(b) $\text{Na}_2\text{SO}_4 \rightarrow$ ionic - dissociate in H_2O
strong.

(c) $\text{H}_2\text{O} \rightarrow$ covalent - no dissociation
nonelectrolyte.

(d) $\text{HCl (aq)} \rightarrow$ dissociate in water b/c it is
~~an acid~~ acid. strong electrolyte



(d) < (a) b/w a & b b > a

a & c. $\text{NaCl} < \underline{\text{Na}_2\text{SO}_4} > \text{LiCl}$