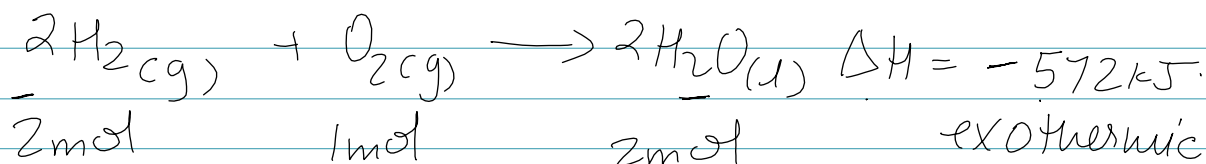
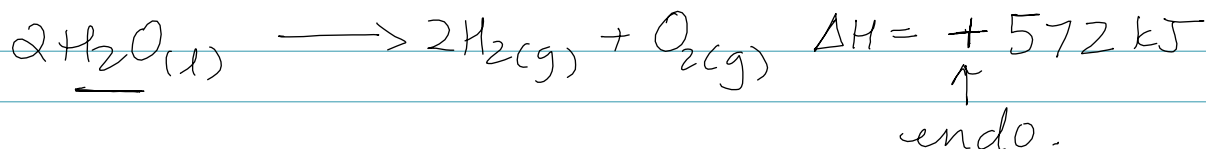


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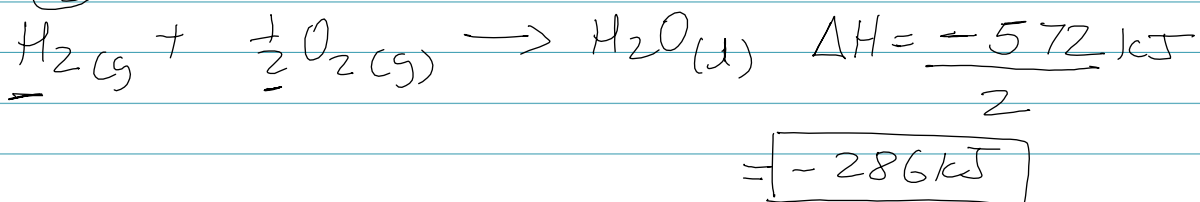
## Manipulating Thermochemical Equations



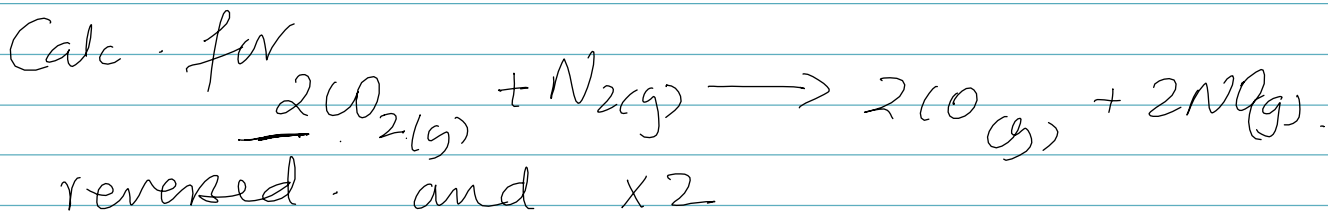
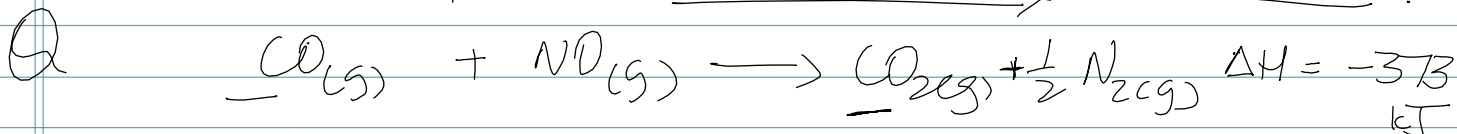
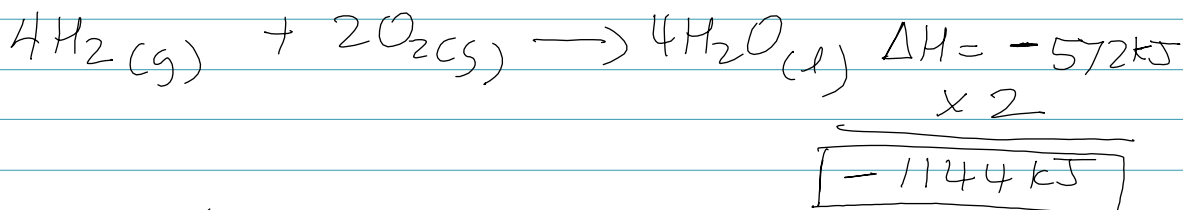
① Reverse eqn.



② Take  $\left(\frac{1}{2}\right)$  the eqn.



③ Double the eqn.



$$\begin{array}{ccc} \Delta H = -373 & & \Delta H = +373 \times 2 \\ & & = \boxed{+746\text{kJ}} \end{array}$$