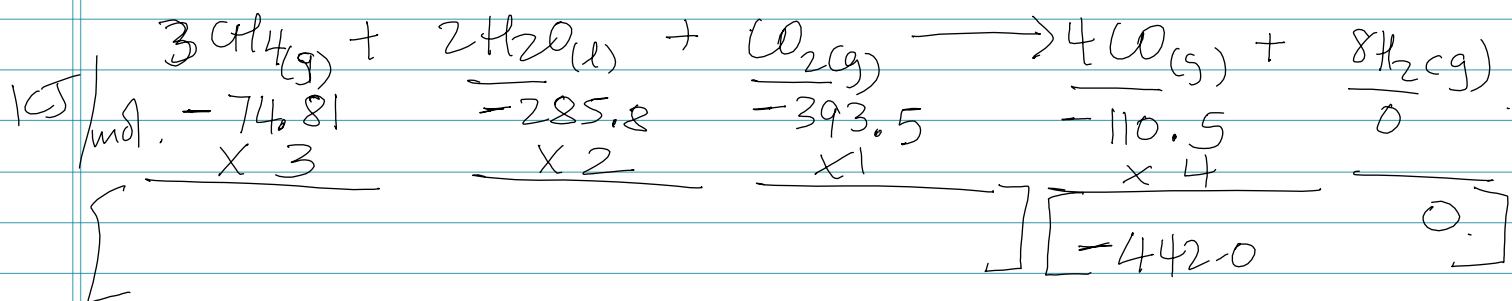


Thermochemistry - ΔH°

1) Use standard ~~etc~~ enthalpies to calculate standard enthalpy change for the following reaction ΔH°



$$\text{products} - \text{reactants} =$$

$$-442.0 \text{ kJ} - (-1189.5 \text{ kJ}) = +747.5 \text{ kJ}$$

$$\sum \text{product} - \sum \text{reactant} = \Delta H^\circ$$

$$\{ \underline{n(\text{CO})} \} - \{ \underline{3 \text{CH}_4} + \underline{2 \text{H}_2\text{O}} + \underline{1 \text{CO}_2} \} \leftarrow$$