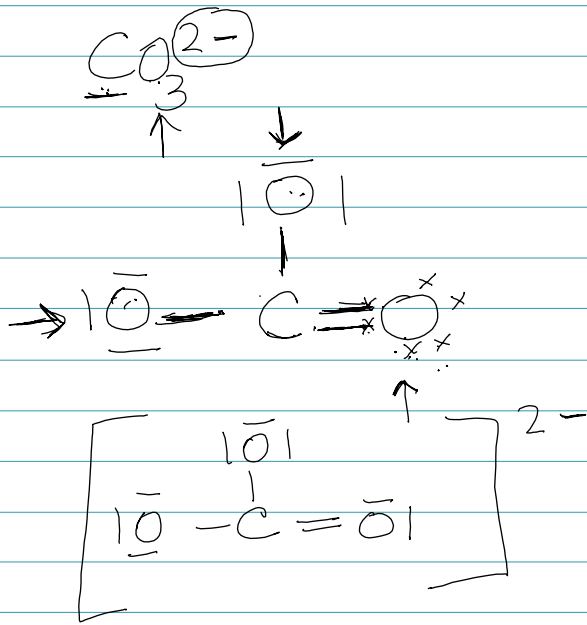


Sapna Singh

Lewis Structure - Polyatomic ions ①



electronic bookkeeping

C has $4e^- = 4$

3(O) has $3 \times 6e^- = 18$

∴ C 2- = 2

24e⁻

3 single bonds - 6e⁻

18e⁻

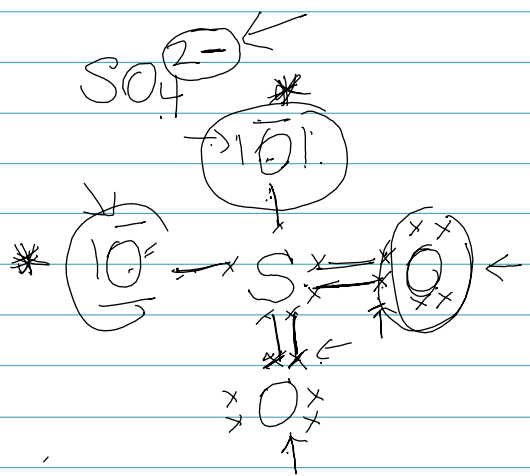
1 more bond as - 2e⁻

16e⁻

4e⁻ on = 0 - 4

12e⁻ / 2

= 6e⁻



S = 6e⁻

4(O) = 24e⁻

2- = 2e⁻

32e⁻

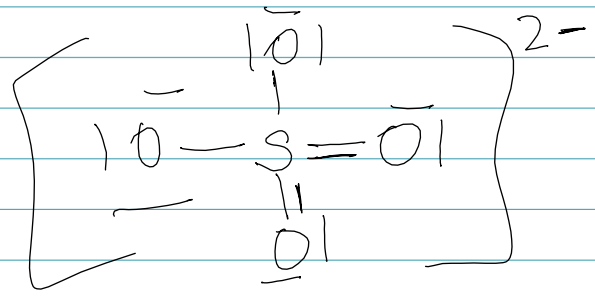
4 single bonds - 8

24e⁻

2 bonds b/w S & O - 4

20e⁻

Sulfur has expanded shell.



filled valence shell for double bonded (O)

4 x 2 = 8

12e⁻ / 2

= 6e⁻