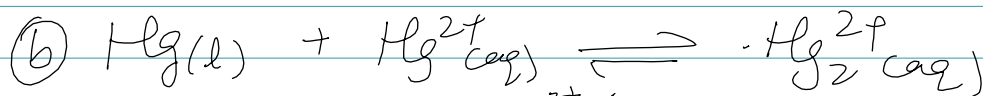


## Eg. ① K<sub>c</sub> expression and Calculation.

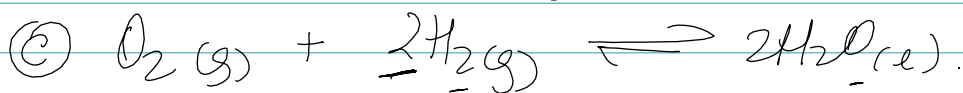
\*) Write the K<sub>c</sub> expression for the following eqn



Ans  $K_c = [\text{CO}_2]$

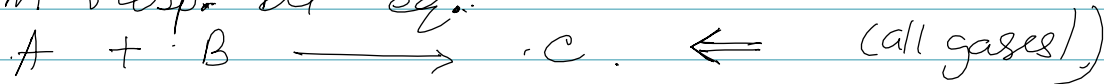


Ans  $K_c = \frac{[\text{Hg}_2^{2+}]}{[\text{Hg}^{2+}]}$



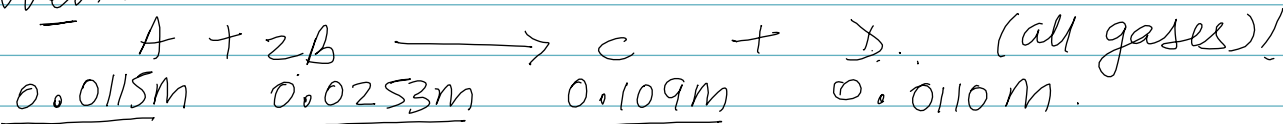
Ans  $K_c = \frac{1}{[\text{O}_2][\text{H}_2]^2}$

② Calculate the K<sub>c</sub> for the reaction below, at 74°C, when the conc. of [A], [B] and [C] are  $1.2 \times 10^{-2} \text{M}$ ,  $0.054 \text{M}$  and  $0.14 \text{M}$  resp. at eq.



$K_c = \frac{[C]}{[A][B]} = \frac{0.14}{(1.2 \times 10^{-2})(0.054)} = \boxed{216}$

③ Determine the K<sub>c</sub> for the reaction below with conc. given.



\*)  $K_c = \frac{[C][D]}{[A][B]^2} = \frac{(0.109)(0.0110)}{(0.0115)(0.0253)^2} = \boxed{163}$