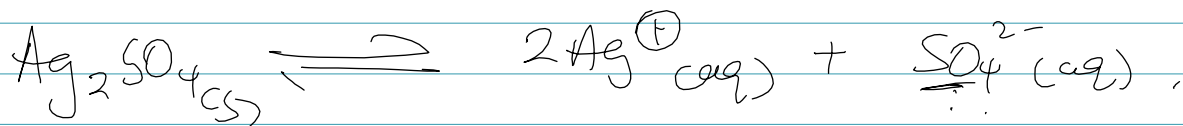


Samarjit

Ksp (3) Common Ion Effect

Calculate the molar solubility of Ag_2SO_4 in 1.00 M Na_2SO_4 . $K_{sp} \text{Ag}_2\text{SO}_4 = 1.4 \times 10^{-5}$.



I	0	1.00 M (from Na_2SO_4)
C	+2s	+s (from Ag_2SO_4)
E	2s	(1+s)

$$1.4 \times 10^{-5} = (2s)^2 (1+s)$$

$$1.4 \times 10^{-5} = 4s^2$$

negligible. ignore!

$$s = 1.9 \times 10^{-3}$$