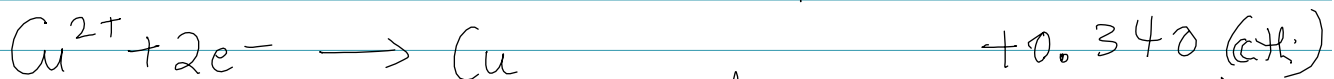
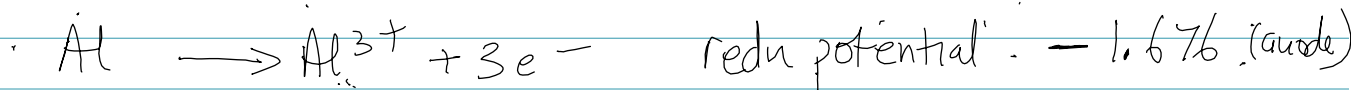
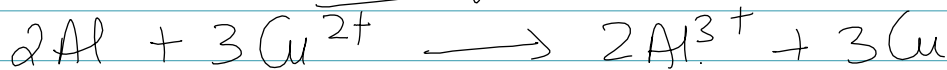


Seyma fupg

Electrochem (5) Calculating E_{cell}° .

① Determine the E_{cell}° for the following reaction.



$$\begin{aligned} E_{cell} &= E_{cath} - E_{anode} \\ &= 0.340 - (-1.676) \\ &= \boxed{2.016V} \end{aligned}$$

2nd way.

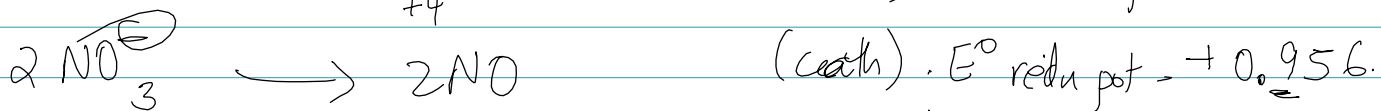
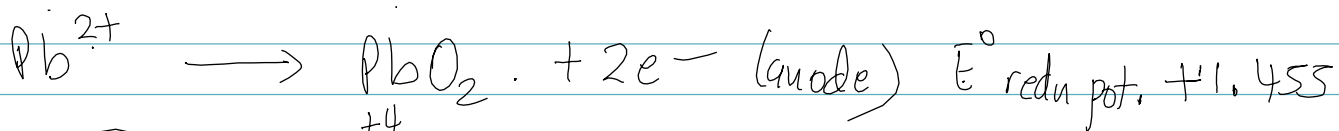
Since Al is oxidized —
~~inv~~ reverse the charge on E°

$$E_{cell} = E_{cath} + E_{anode}$$

$$= 0.340 + 1.676$$

$$= 2.016$$

② Determine E_{cell}° for:



$$\begin{aligned} E_{cell} &= E_{cath} - E_{anode} \\ &= 0.956 - (+1.455) \end{aligned}$$

$$= \boxed{-0.499V}$$

2nd, inv- anode

$$-1.455$$

$$+0.956$$

$$= \boxed{-0.499V}$$