

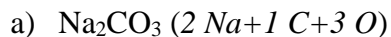
Ch3/ PowerPoint Study-1 Molar Mass/Moles/Atoms Name:

Answer these questions as you are watching the videos. They are due in class.

These questions are not just for you to answer but also to prepare you for the exam.

*Make sure you understand what you are writing and not just copy from the text book. **Show all work.***

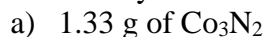
1) Calculate the molar mass of the following compounds:



c) copper (I) arsenide

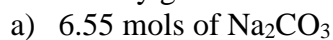
d) diphosphorous pentaoxide

2) How many mols are in the following (you can use the molar masses from question 1):



b) 2.45 g of copper (I) arsenide

3) How many grams are in the following:



b) 6.55×10^{22} molecules of potassium oxide

- 4) How many molecules are in the following:
- a) 10.2 moles of diphosphorous pentaoxide

 - b) 9.66 grams of potassium oxide
- 5) How many oxygen atoms are in 3.55 g of Na_2CO_3 .
- 6) What is mass percent of each element in copper (I) carbonate?
- 7) What is the mass percent of water in magnesium sulfate heptahydrate?
- 8) What is the empirical formula of a compound containing 77.7% iron and 22.3 % oxygen?