

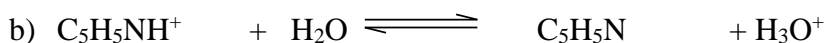
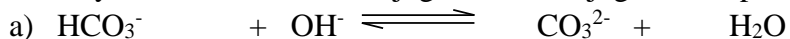
1) Write the conjugate base for the following acids:



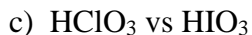
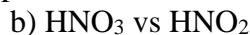
2) Write the conjugate acid for the following bases:



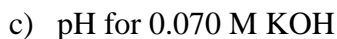
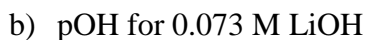
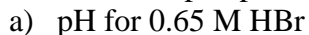
3) Identify the acid-base and conjugate acid/conjugate base pair in the following equations.



4) Circle the stronger acid in the following pairs:



5) Calculate the pH, pOH,  $[H_3O^+]$  or  $[OH^-]$  concentrations as indicated in the problem:



6) Classify the following as acidic, basic or neutral salts:

a) LiCl

b) KNO<sub>3</sub>

c) NH<sub>4</sub>Cl

d) Na<sub>3</sub>PO<sub>4</sub>

7) Identify the Lewis acid and the base in the following reactions.

