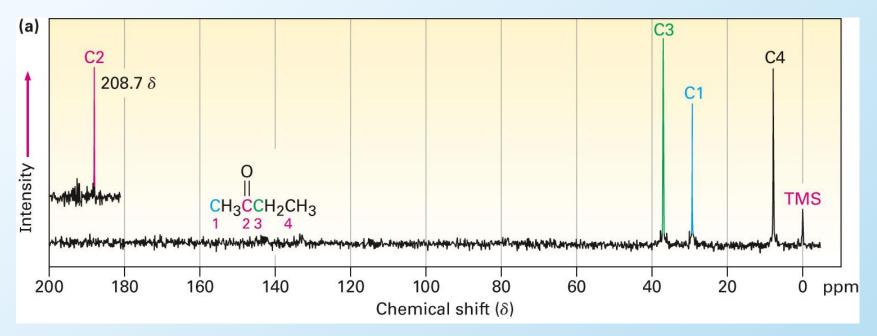
Nuclear Magnetic Resonance Part 3 Carbon - 13

Dr. Sapna Gupta

Carbon-13

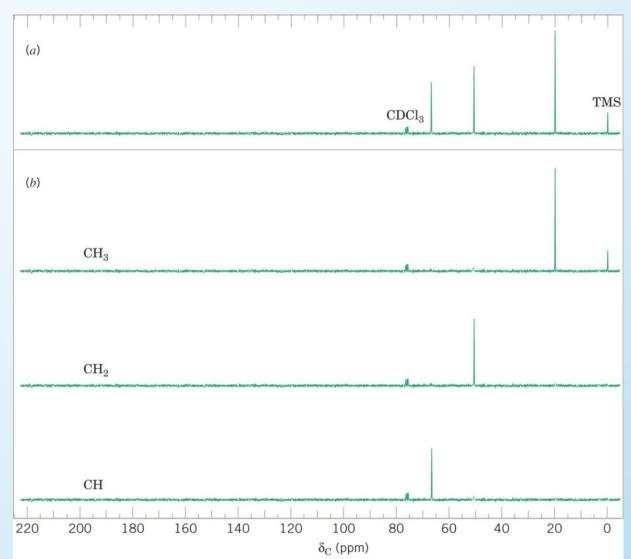
- ¹²C has no magnetic spin.
- ¹³C has a magnetic spin, but is only 1% of the carbon in a sample.
- Signals are weak, getting lost in noise.
- Hundreds of spectra are taken, averaged.
- Below is the spectrum for **2- butanone**



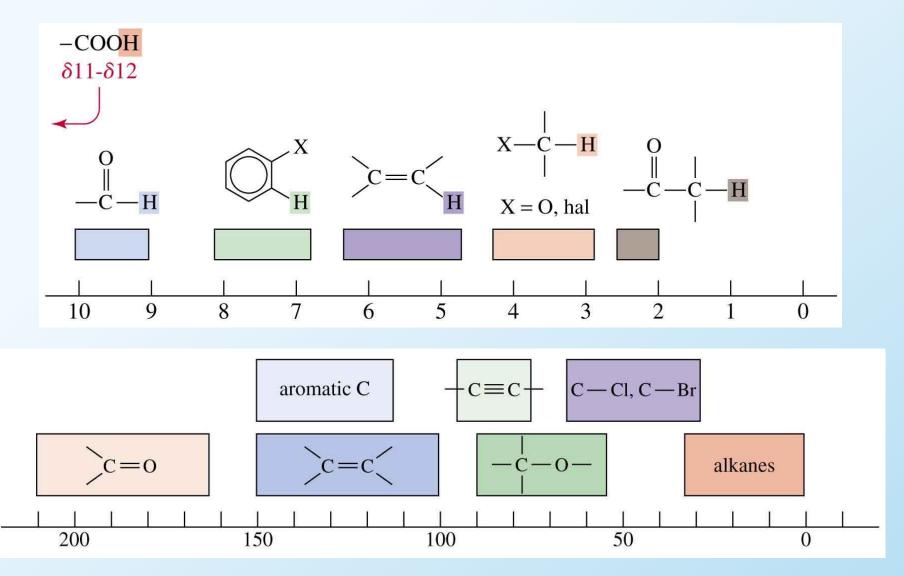
DEPT for C-13

Example: 1-chloro-2-propanol

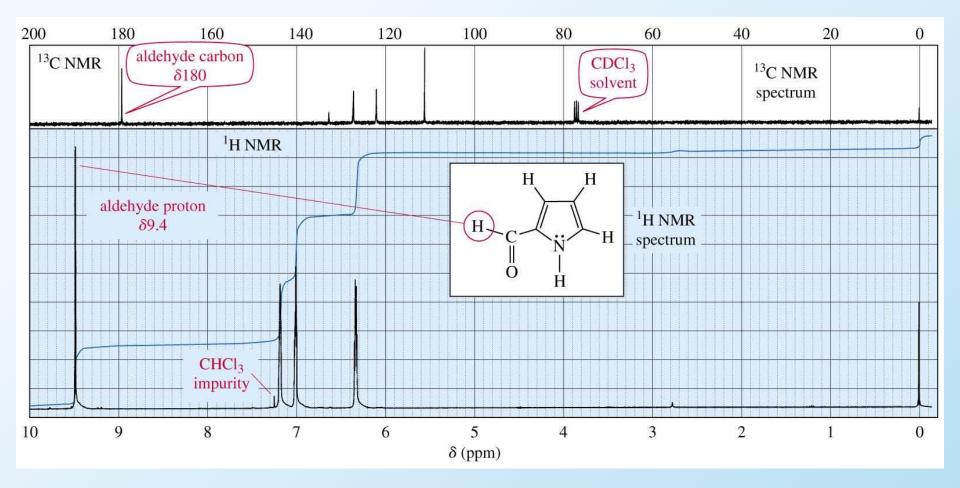
- (a) The broadband decoupled spectrum and
- (b) a set of DEPT spectra showing the separate CH, CH_2 , and CH_3 signals



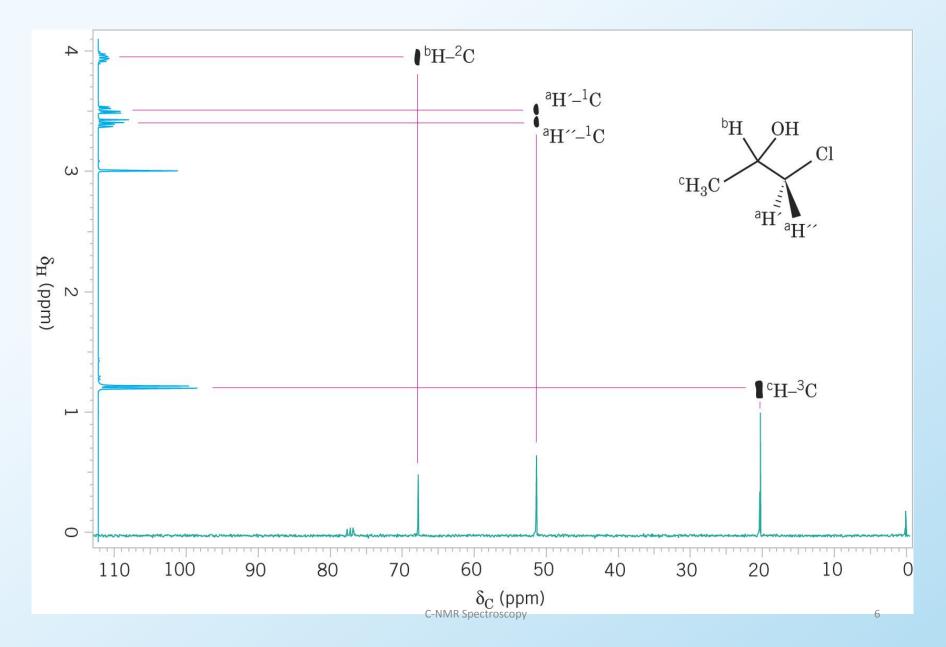
Hydrogen and Carbon Chemical Shifts



Combined ¹³C and ¹H Spectra



H and C coupled NMR (HETCOR)



Spin-Spin Splitting

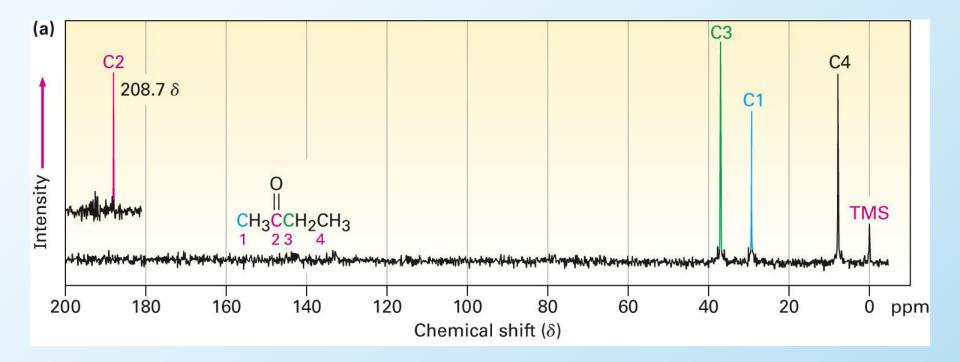
- It is unlikely that a ¹³C would be adjacent to another ¹³C, so splitting by carbon is negligible.
- ¹³C <u>will</u> magnetically couple with attached protons and adjacent protons.
- These complex splitting patterns are difficult to interpret.

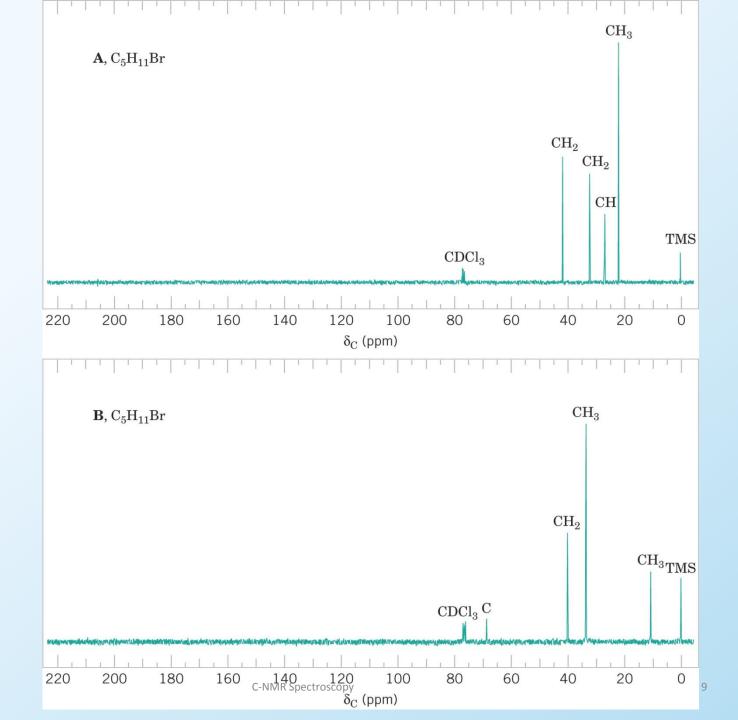
Interpreting ¹³C NMR

- The number of different signals indicates the number of different kinds of carbon.
- The location (chemical shift) indicates the type of functional group.
- The peak area indicates the numbers of carbons (if integrated).
- The splitting pattern of off-resonance decoupled spectrum indicates the number of protons attached to the carbon.

C-13 NMR of 2-butanone

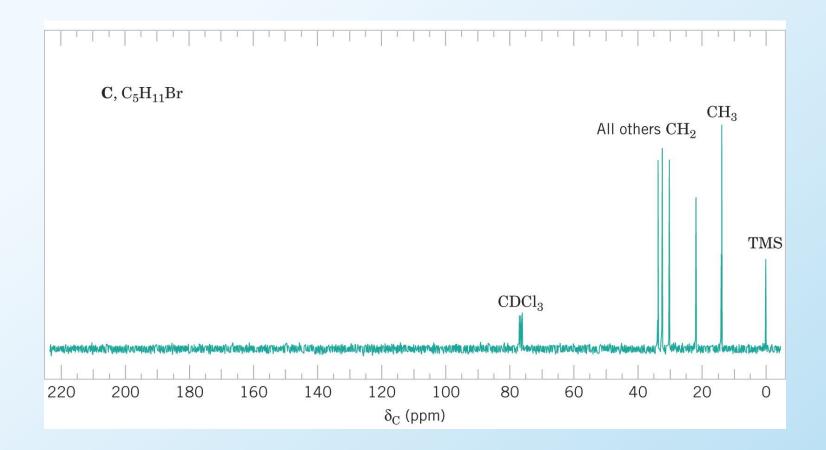
- CH₃COCH₂CH₃
- Shows the carbon of C=O (carbonyl) group on the left, and the alkyl carbons on the right – upfield.





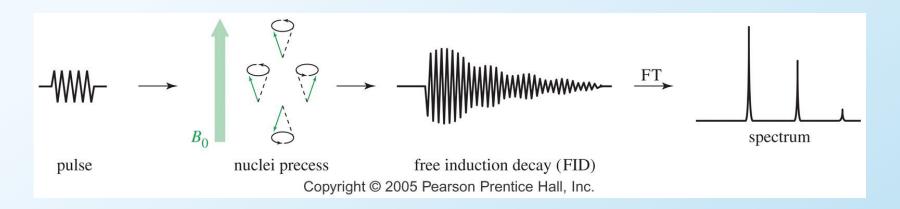
Problems

Problem



Fourier Transform NMR

- Radio-frequency pulse given.
- Nuclei absorb energy and precess (spin) like little tops.
- A complex signal is produced, then decays as the nuclei lose energy.
- Free induction decay is converted to spectrum.



Key Concepts

- Chemical shifts
- Interpret a C-13 NMR
- Predict signals for C-13 NMR for a compound