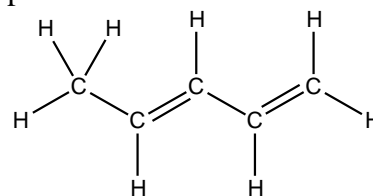
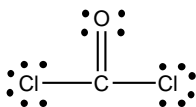
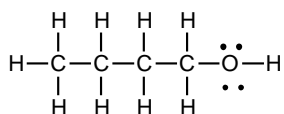


Power Point Study– 04-1 Intermolecular Forces **Name:** _____

- Which force is stronger – hydrogen bond or dipole-dipole?
- Circle the element with the higher electronegativity in the following pairs.
 - F or Br
 - N or O
 - Se or S
- Indicate, using the dipole moment arrow, the direction of the dipole moment in the following molecules. (Hint: draw the 3D structure of the molecule first)
 - SO₂
 - CHCl₃
 - SO₃

- What is the strongest intermolecular force in the following compounds?

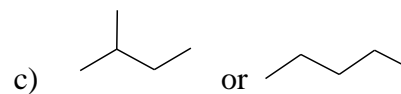
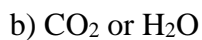
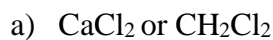


- In which type of compounds are the intermolecular higher: branched or unbranched compounds of the same molecular formula?

- Circle the compound with the higher dipole moment in the following pairs.

CO ₂ or COS	HF or HCl	CH ₃ Cl or CH ₂ Cl ₂	HOH or CH ₃ OCH ₃
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- Circle the compound with the higher intermolecular force.



- Label the compound with the lowest and the highest intermolecular force in the following sets of compounds.

