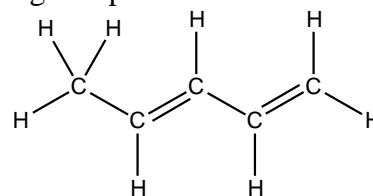
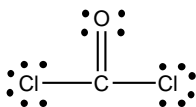
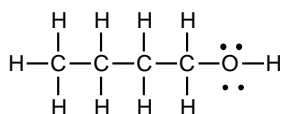


Power Point Study– 01-4 Intermolecular Forces **Name:** _____

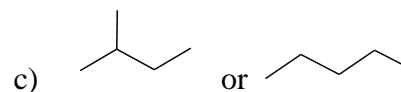
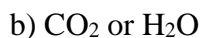
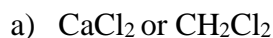
- Which force is stronger – hydrogen bond or dipole-dipole?
- Circle the element with the higher electronegativity in the following pairs.
 - F or Br
 - N or O
 - Se or S
- Indicate, using the dipole moment arrow, the direction of the dipole moment in the following molecules. (*Hint: draw the 3D structure of the molecule first*)
 - SO₂
 - CHCl₃
 - SO₃

- What is the strongest intermolecular force in each of the following compounds?



- In which type of compounds are the intermolecular **higher**: branched or unbranched compounds of the same molecular formula?
- Circle the compound with the **higher** dipole moment in the following pairs.

CO ₂ or COS	HF or HCl	CH ₃ Cl or CH ₂ Cl ₂	HOH or CH ₃ OCH ₃
------------------------	-----------	---	---
- Circle the compound with the **higher** intermolecular force.



- Label the compound with the **lowest** and the **highest** intermolecular force in the following sets of compounds.

