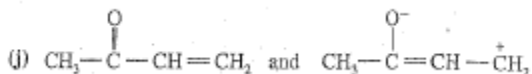
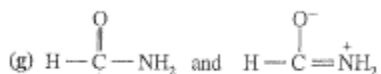
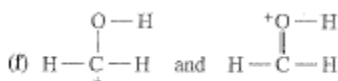
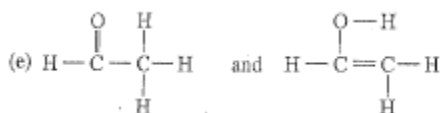
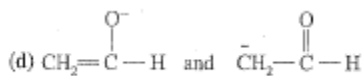
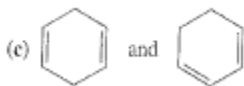
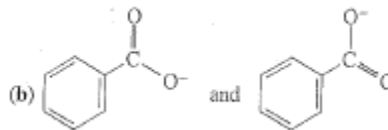
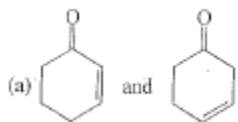
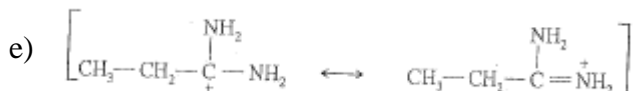
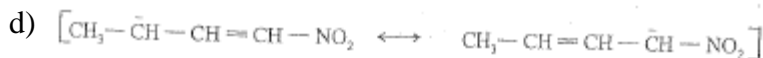
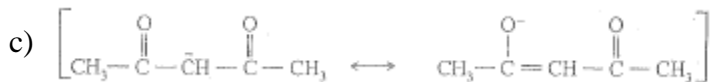
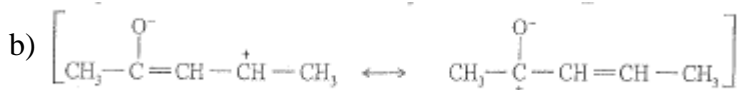


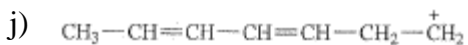
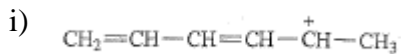
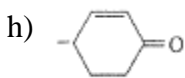
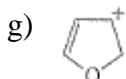
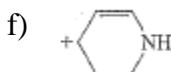
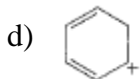
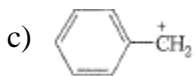
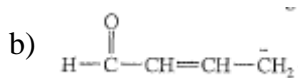
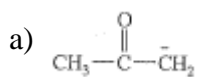
- 1) Determine whether the following pairs of structures are different compounds or resonance structures of the same compound.



- 2) In the following sets of resonance structures, label the major and minor contributors and state which structures would be of equal energy. Add any missing resonance forms.



3) Draw the important resonance forms to show the delocalization of charges in the following ions.



4) For each pair of ions determine which is more stable. Use resonance to explain your answer.

