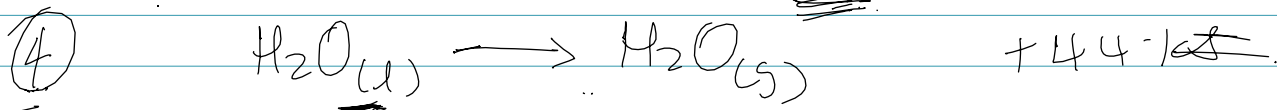
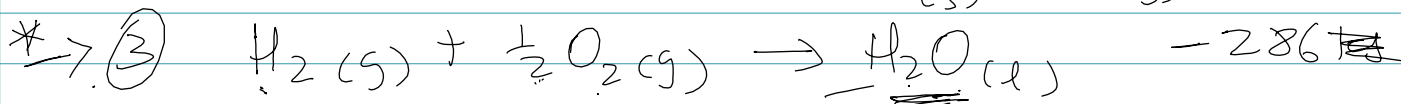
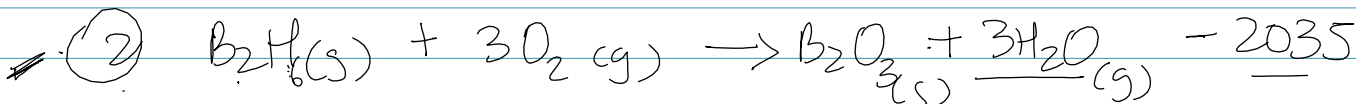
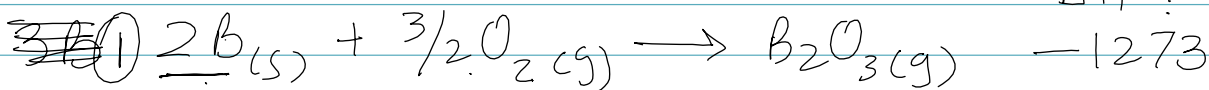


Hess's Law (2)

Find ΔH for $2B(s) + 3H_2(g) \rightarrow B_2H_6(g)$ $\Delta H = ?$
 given



Ans (a) B is in eq(1) & 2mols. eq(1) as is.

(b) H_2 is in eq(3) but 1mol. eq(3) $\times 3$.

(c) B_2H_6 is in eq(2) & 1mol. but reactant! eq(2) ~~as is~~ reverse to get B_2H_6 as product.

(d) eq(2) ~~with~~ has $3H_2O(g)$ so eq(4) $\times 3$ & eq(3) has $3H_2O(l)$

