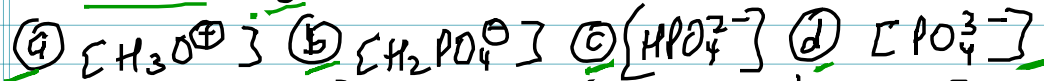
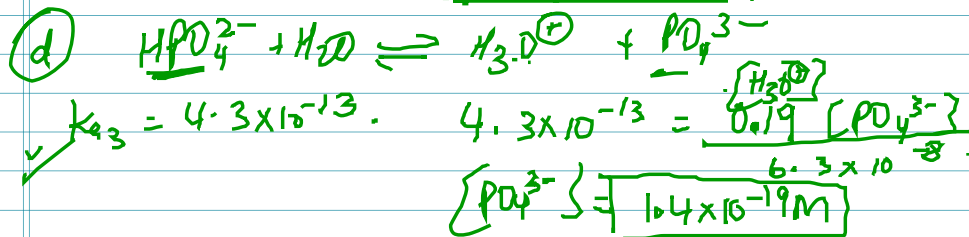
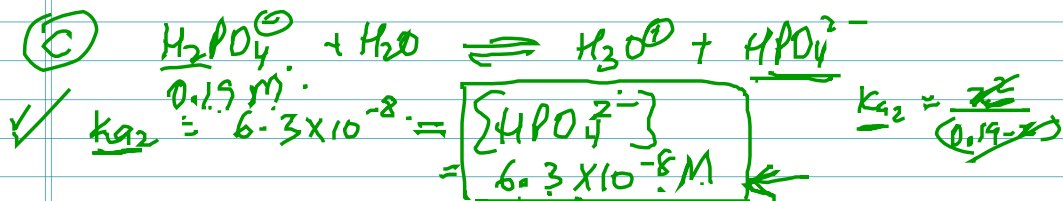
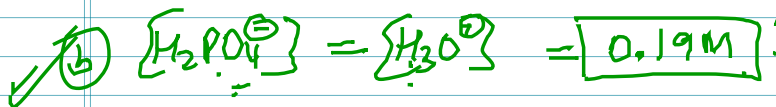
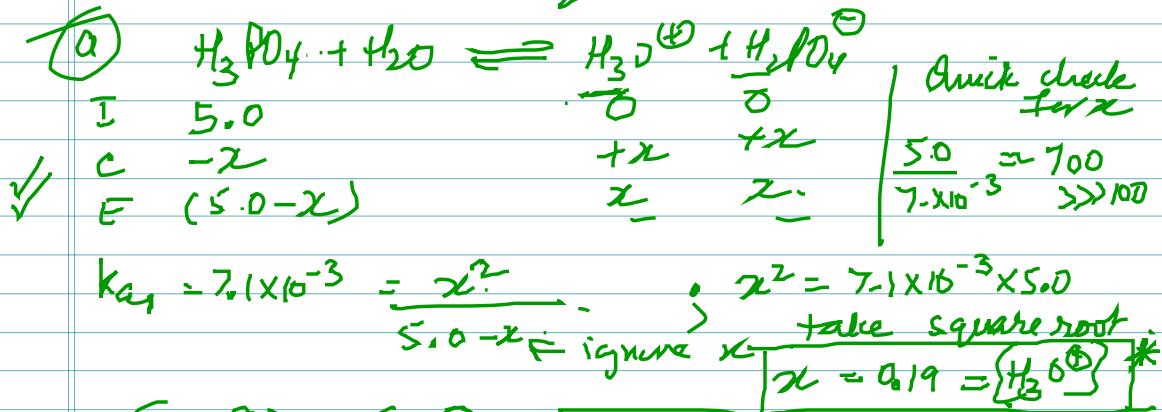


Sapna Syors --

Acid-Base Eq (5) pH of Polyprotic Acid ①Calculate the conc. of the following ions in a 5.0M H_3PO_4 soln.

$$K_{a1} = 7.1 \times 10^{-3}, K_{a2} = 6.3 \times 10^{-8}, K_{a3} = 4.3 \times 10^{-13}$$

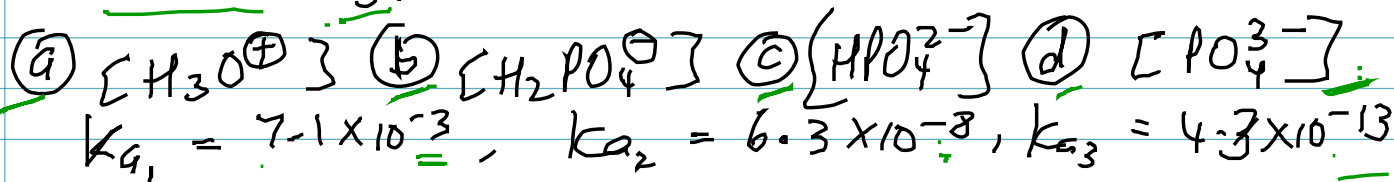
$K_{a1} > K_{a2}$ by 10^5 so most ionization
 $[\text{H}_3\text{O}^+]$ in from 1st ionization.



Sapna Syptis --

Acid-base Eq (5) pH of Polyprotic Acid ①

Calculate the conc. of the following ions in a 5.0M H_3PO_4 soln.



$K_{a1} > K_{a2}$ by 10^5 so most ionization $[H_3O^+]$ in first ionization.

