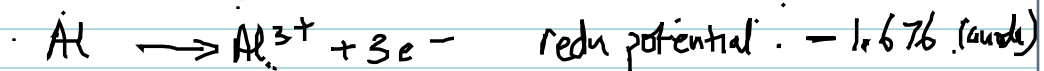
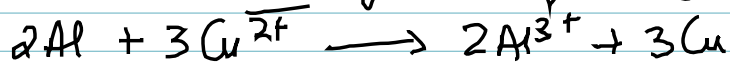


Summary

Electrochem (5) Calculating E_{cell}° .① Determine the E_{cell}° for the following reaction.

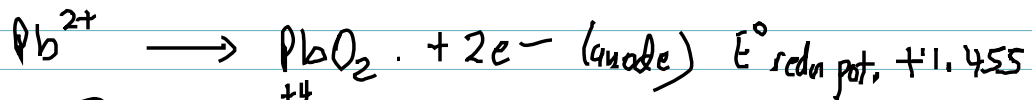
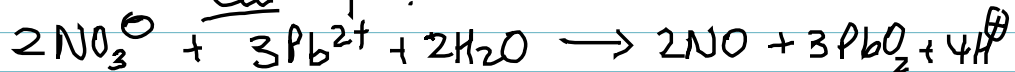
$$\begin{aligned} E_{cell} &= E_{cath} - E_{anode} \\ &= 0.340 - (-1.676) \\ &= \boxed{2.016V} \end{aligned}$$

2nd way.

Since Al is oxidized —
invert reverse the charge on

$$E^{\circ} + 1.676$$

$$\begin{aligned} E_{cell} &= E_{cath} + E_{anode} \\ &= 0.340 + 1.676 \\ &= 2.016 \end{aligned}$$

② Determine E_{cell}° for:

$$\begin{aligned} E_{cell} &= E_{cath} - E_{anode} \\ &= 0.956 - (+1.455) \\ &= \boxed{-0.499V} \end{aligned}$$

2nd,
inv. anode

$$-1.455$$

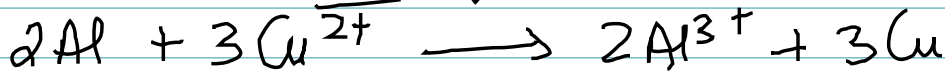
$$+0.956$$

$$\boxed{-0.499V}$$

Summary

Electrochem (5) Calculating E°_{cell} .

① Determine the E°_{cell} for the following reaction.



$$\begin{aligned} E_{\text{cell}} &= E_{\text{cath}} - E_{\text{anode}} \\ &= 0.340 - (-1.676) \\ &= \boxed{2.016\text{V}} \end{aligned}$$

2nd way.

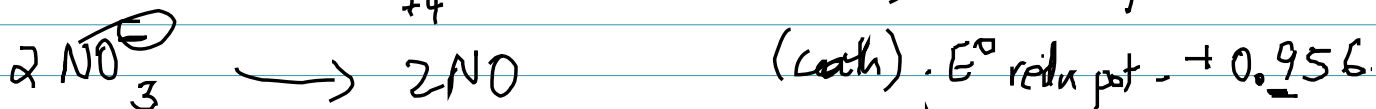
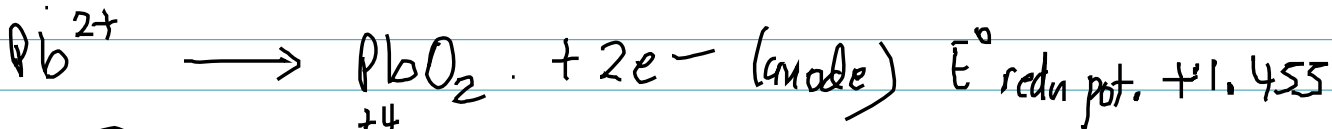
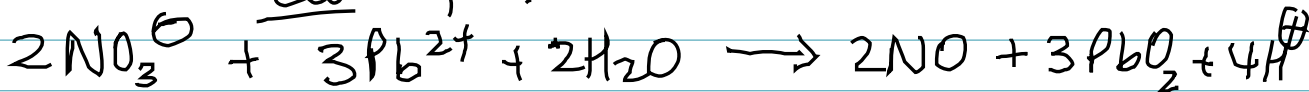
Since Al is oxidized -
~~reverse~~ reverse the charge on
 E° $+1.676$

$$E_{\text{cell}} = E_{\text{cath}} + E_{\text{anode}}$$

$$= 0.340 + 1.676$$

$$= 2.016$$

② Determine E°_{cell} for:



$$E_{\text{cell}} = E_{\text{cath}} - E_{\text{anode}}$$

$$= 0.956 - (+1.455)$$

$$= \boxed{-0.499\text{V}}$$

2nd
inv- anode

$$-1.455$$

$$+0.956$$

$$\boxed{-0.499\text{V}}$$